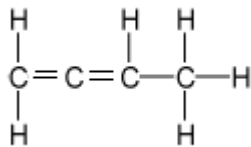
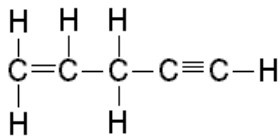


Pojmenuj sloučeninu/sestav strukturní vzorec:



a)

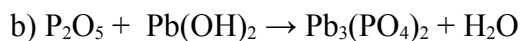
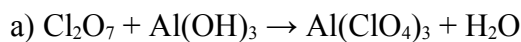
b) hex-1-en



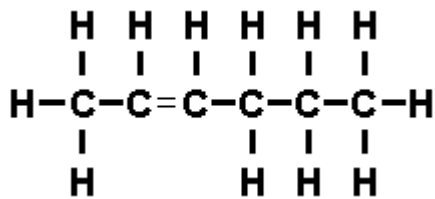
c)

d) hepta-1-en-3,5-diyn

Domácí úkol: Pojmenuj reaktanty a produkty, rovnice vyčísli:

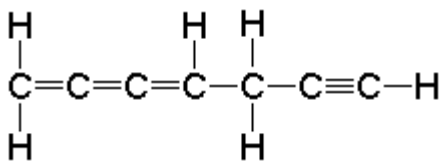


Napiš název sloučeniny/sestav strukturní vzorec:



a)

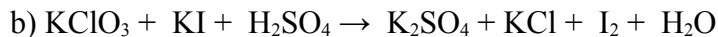
b) penta-1,4-diyn



c)

d) penta-1-en-3-yn

Domácí úkol:

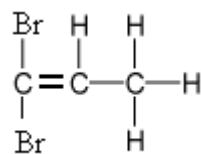


1. Vypočti kolik látky musíš rozpustit a v jakém množství vody, abys připravil 1450 g roztoku s koncentrací 40 %.

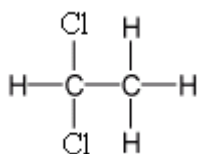
2. Ve vodě o hmotnosti 1440 g jsme rozpustili 100 g látky. Jakou koncentraci má roztok, který vznikl?

3. Kolik látky musíme rozpustit ve 820 g vody, aby vznikl 38% roztok?

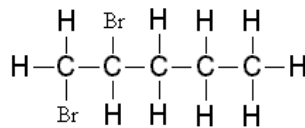
4. V jakém množství vody musíme rozpustit 100 g látky, aby vznikl 18% roztok?



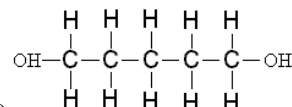
a)



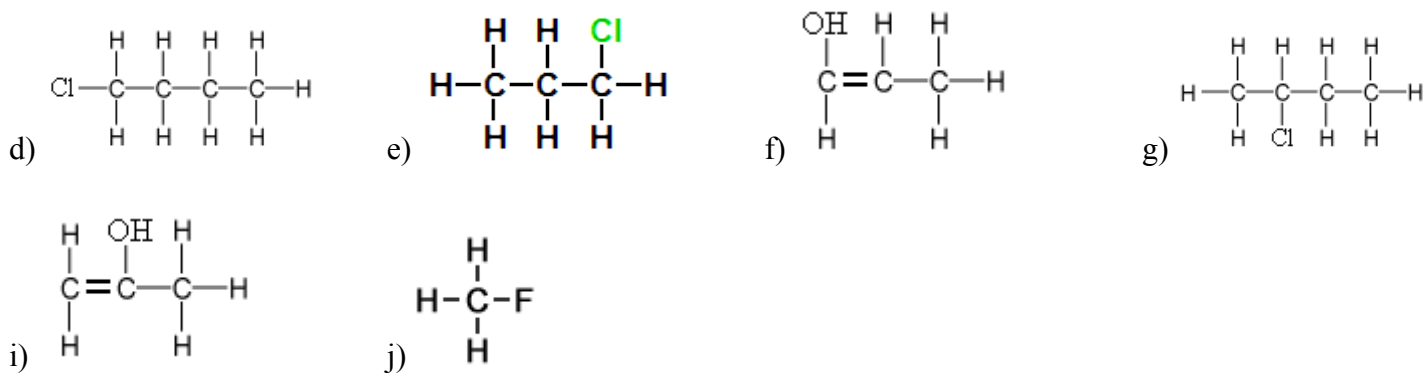
b)



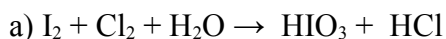
c)



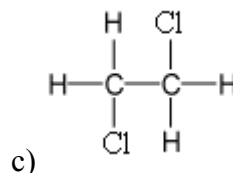
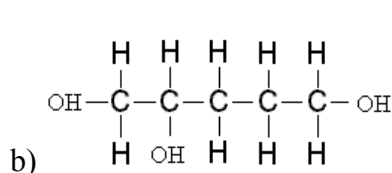
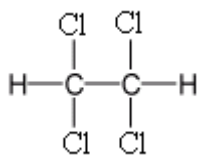
c)



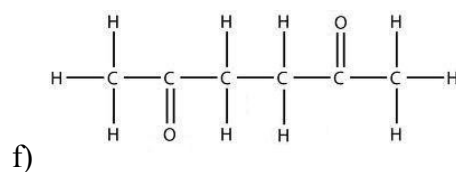
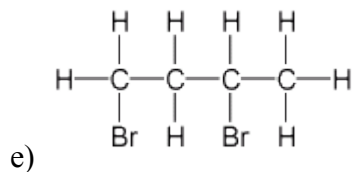
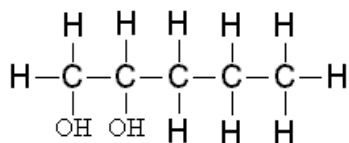
Domáci úkol: Pojmenuj reaktanty a produkty, rovnice vyčísli:



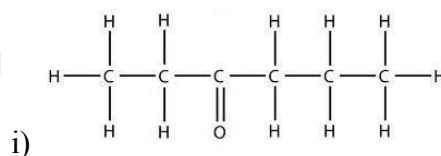
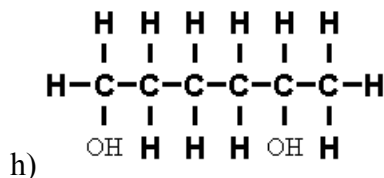
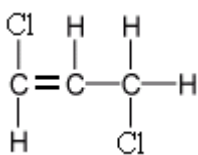
a)



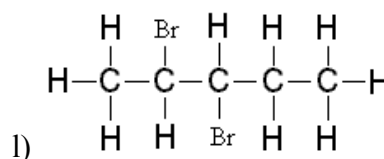
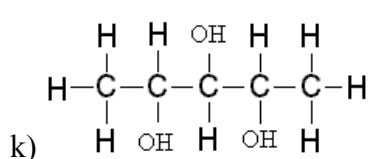
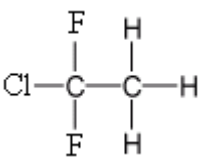
d)



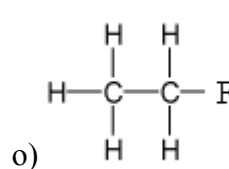
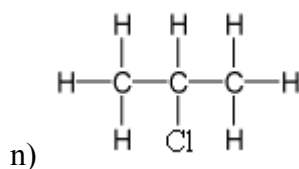
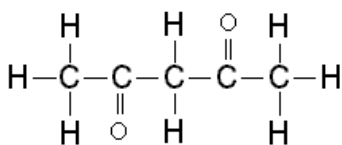
g)



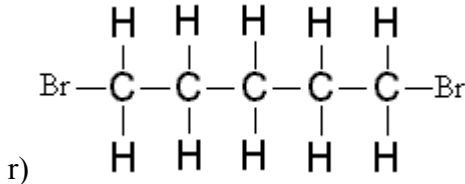
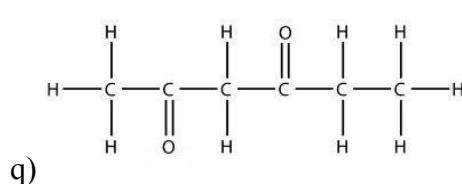
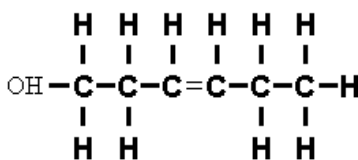
j)

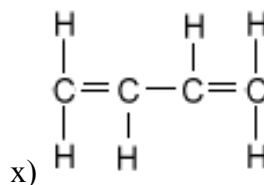
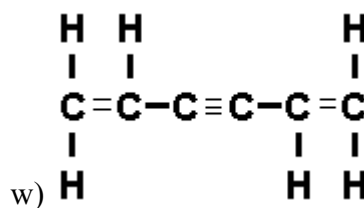
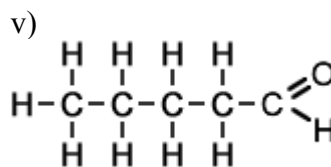
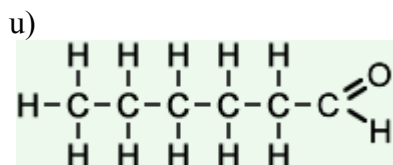
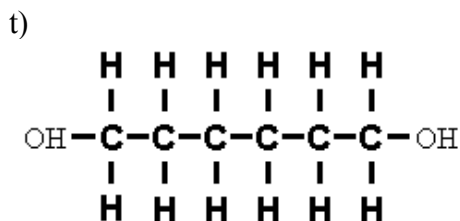
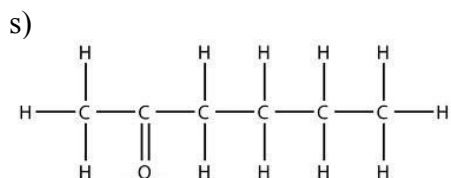


m)

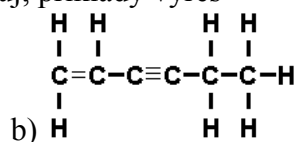
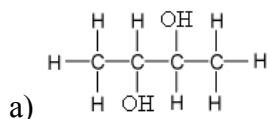


p)

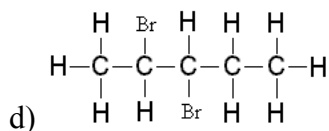
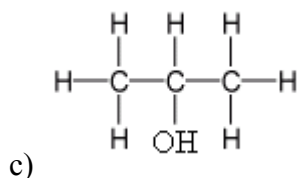




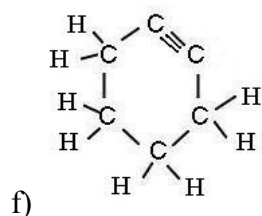
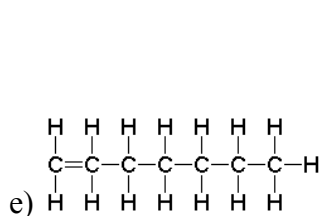
Uvedené sloučeniny pojmenuj, příklady vyřeš



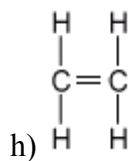
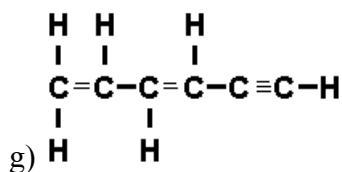
1. Kolik gramů chloridu železnatého vznikne reakcí 250 g sulfidu železnatého s kyselinou chlorovodíkovou? (sulfid železnatý reaguje s kyselinou chlorovodíkovou za vzniku chloridu železnatého a sulfanu)



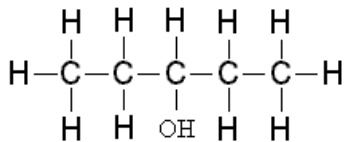
2. Kolik gramů chloridu stříbrného vznikne reakcí 20 g chloridu sodného s dusičnanem stříbrným? (chlorid sodný reaguje s dusičnanem stříbrným za vzniku dusičnanu sodného a chloridu stříbrného)



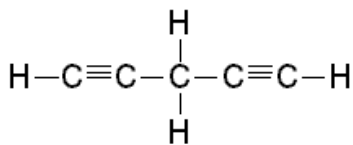
3. Pojmenuj reaktanty a produkty, rovnici vyčísli:  
 $\text{As}_2\text{O}_3 + \text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{AsO}_4 + \text{NO}$



4. Kolik g zinku musí reagovat s kyselinou chlorovodíkovou, aby vzniklo 8 g vodíku?



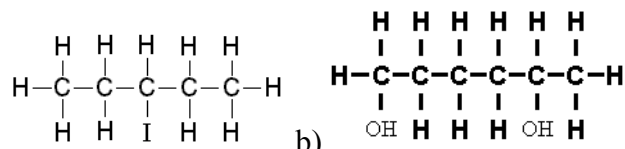
i)



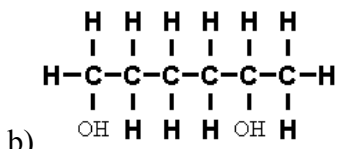
j)

5. Kolik gramů hydroxidu sodného vznikne reakcí 5 g sodíku s vodou?

0. Máte 200 g 20% roztoku, ale potřebujete jen 100 g 15% roztoku. Jak to zařídíte?

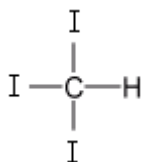


a)

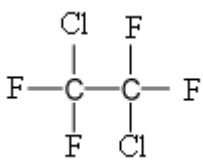


b)

1. Vypočítej hmotnost uhličitanu vápenatého, kterou potřebujeme k výrobě 112 kg páleného vápna. ( $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ )

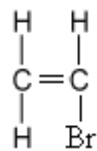


c)

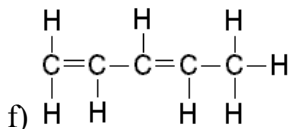


d)

2. Kolik gramů síranu sodného vznikne při reakci 250 g hydroxidu sodného s kyselinou sírovou?

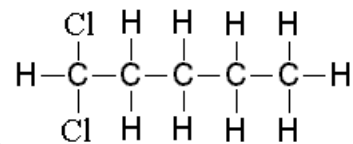


e)

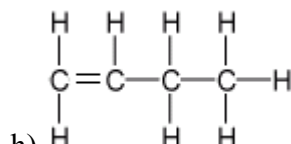


f)

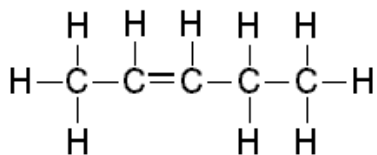
3. Kolik gramů chloridu sodného vznikne při neutralizaci 22 g kyseliny chlorovodíkové?



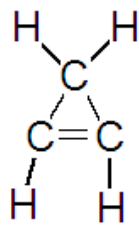
g)



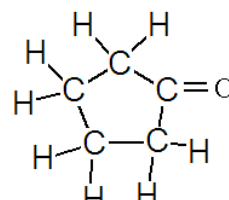
h)



i)

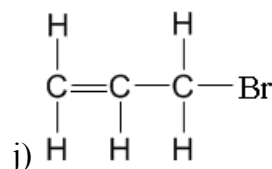
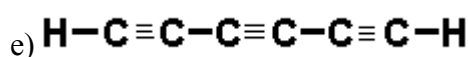
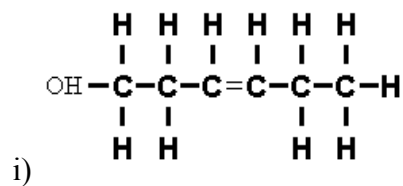
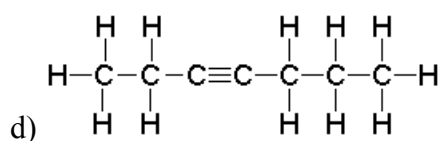
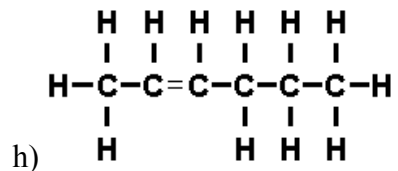
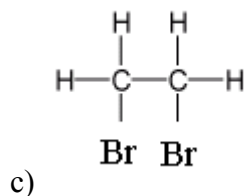
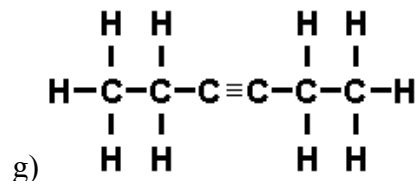
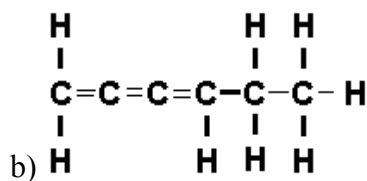
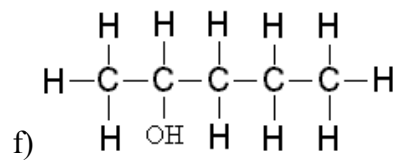
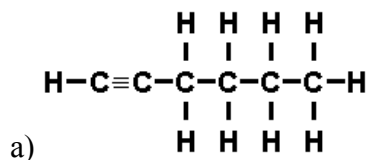


j)



k)

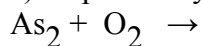
Máte k dispozici 300 gramů 30% roztoku. Vy ale potřebujete jen 100 gramů 18% roztoku. Jak to zařídíte? Kolik roztoku odeberete a jaké množství vody přidáte?



1. Co je to krakování?
2. Železo se ve vysoké peci získává redukcí \*\*\* uhlíkem. Co se skrývá za hvězdičkami?
3. Copak se děje s oxidačním číslem prvku, který se při reakci redukuje?
4. Které látky se označují jako ketony?
5. O které reakci se řekne, že je exotermická?
6. Která látka obsahuje dva atomy uhlíku, pět atomů vodíku a jednu hydroxylovou skupinu?

1. Reakcí arsenu a kyslíku vznikne oxid arsenitý

a) Doplňte a vyčíslete rovnici



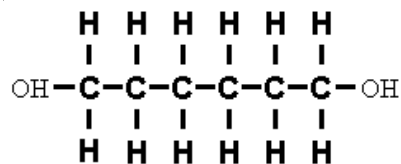
b) Kolik arsenu potřebujeme, chceme-li reakcí získat 100 gramů oxidu arsenitého?

2. Reakcí kyseliny chlorovodíkové a hliníku vznikne \*\*\* a uvolní se \*\*\*

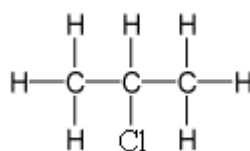
a) Zapište a vyčíslete rovnici:

b) Kolik kyseliny chlorovodíkové musíme přidat ke 20 g hliníku, aby se všechny reaktanty spotřebovaly.

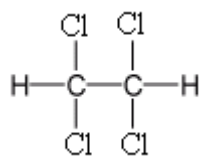
Pojmenuj uvedené sloučeniny:



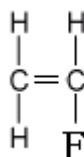
a)



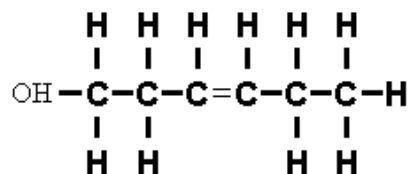
e)



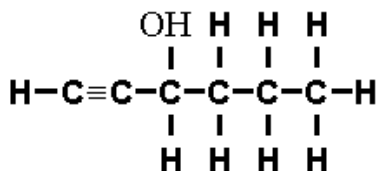
b)



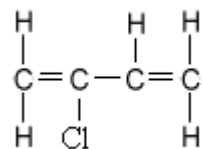
f)



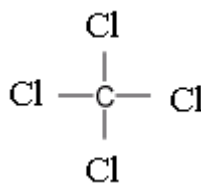
c)



g)



d)



h)

Domácí úkol:

1. Zapiš rovnici a vyčíslí. Amoniak reaguje s chlornanem sodným za vzniku dusíku, chloridu sodného a vody.

2. Pokud chceme uvedenou reakcí připravit 70 g chloridu sodného, kolik musí do reakce vstoupit reaktantů?

Síran sodný reaguje s chloridem vápenatým za vzniku \*\*\* vápenatého a chloridu \*\*\*.

a) Sestav rovnici a vyčíslí stechiometrické koeficienty.

b) Kolik chloridu vznikne, pokud do reakce vstoupí 40 gramů síranu sodného?

a) chlorid měďnatý

b)  $\text{Li}_2\text{CO}_3$

c) jodid cínatý

d)  $\text{SnCl}_2$

e) síran zinečnatý

f)  $\text{VO}_2$

g) chlorid lithný

h)  $\text{PbHPO}_4$

i) hydrogensířičitan manganatý

j)  $\text{B}_2\text{O}_3$

k) oxid fosforičitý

l)  $\text{KNO}_3$

m) chlorid draselný

n)  $\text{Br}_2\text{O}_7$

o) bromid draselný

p)  $\text{Al}(\text{H}_2\text{PO}_4)_3$

q) oxid draselný

r)  $\text{H}_3\text{PO}_2$

- s) oxid sirmatý
- t)  $\text{LiNO}_3$
- u) sulfid železitý
- v)  $\text{WO}_2$
- w) siřičitan sodný
- x)  $\text{PbO}$
- y) hydroxid barnatý
- z)  $\text{LiF}$